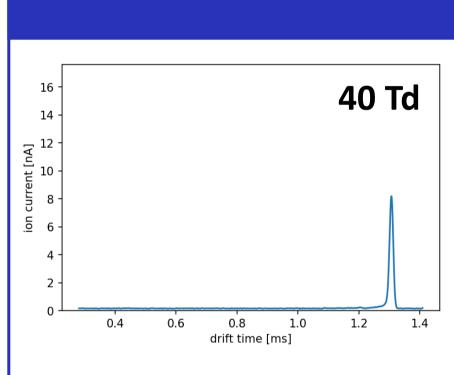
# Investigations on the field dependence of cluster dynamics using High Kinetic Energy IMS (HiKE-IMS)

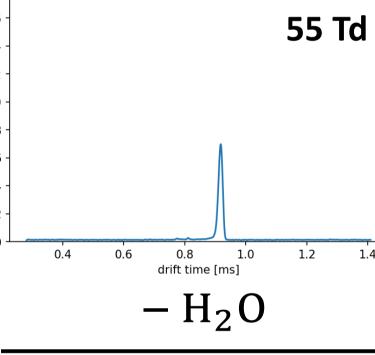
# <u>Florian Stappert<sup>1</sup></u>; Duygu Erdogdu<sup>1</sup>; Maria Allers<sup>2</sup>; Ansgar T. Kirk<sup>2</sup>; Walter Wissdorf<sup>1</sup>; Hendrik Kersten<sup>1</sup>; Stefan Zimmermann<sup>2</sup>; Thorsten Benter<sup>1</sup>

# Introduction

The dynamic clustering and declustering processes of ions with separation with DMS is based on the different mobilities of ions neutrals in the gas phase are very important in many modern in an oscillating electric field.<sup>[3]</sup> In previous results, the dynamic methods used in chemical analytics. The electrical mobility K cluster system with neutrals, like added modifiers, was which governs ion separation in Ion Mobility Spectrometry identified as one root cause of this effect in certain cases.<sup>[4]</sup> (IMS) is dependent on the collision cross section which is Thus, a deeper understanding of cluster stabilities and the generally influenced by clustering reactions<sup>[1]</sup> and therefore also resulting dynamics at high reduced field strength is essential for by the addition of chemical modifiers to the gas phase.<sup>[2]</sup> the understanding of these complex chemical systems.

In addition to this technique, which is running at a red. field By using a High Kinetic Energy IMS (HiKE-IMS), similar high field strength below 10 Td, clustering processes under high field conditions can be recreated with a less complex analytical conditions are of great relevance, too. Typical regions with high method. Based on a classic drift tube IMS, a reduced field field conditions are the transfer stages of mass spectrometers strength up to >120 Td can be realized by running the system at and **D**ifferential **M**obility **S**pectrometry (DMS) systems. The ion reduced pressure of 20–30 mbar (see "Experimental").<sup>[5]</sup>



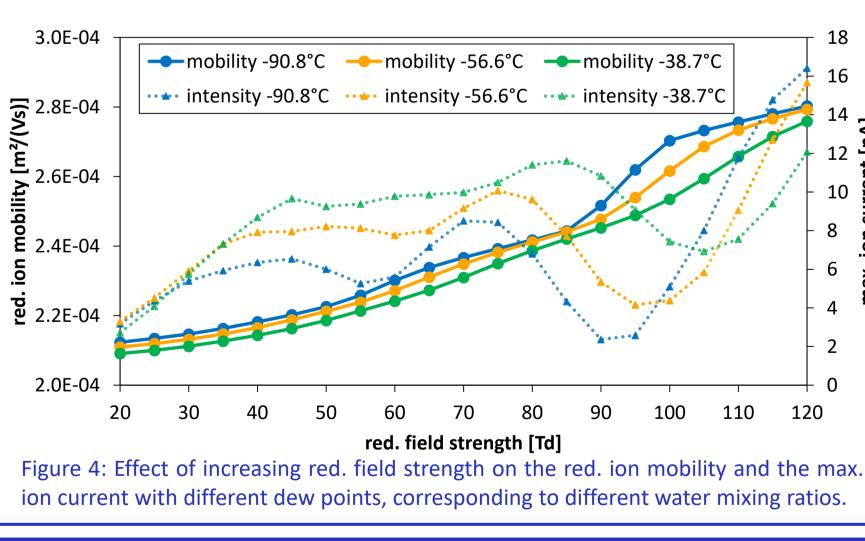


### drift time [ms] $[H+(H_2O)_2]^+$

### $[H+(H_2O)_3]^+$ Figure 2: Experimental HiKE-IMS-spectra in dependence of representative reduced field strength (from left: 40 Td, 55 Td,

signal can be observed which shifts to lower combination of two cluster species. strength (fig. 2). The cluster transitions times.

Without analytes present, a dominant are observed as temporary broadening of pure water clusters ("RIP") of the peaks as the result of a drift times with increasing red. field Note the pre-RIP-signals at lower drift

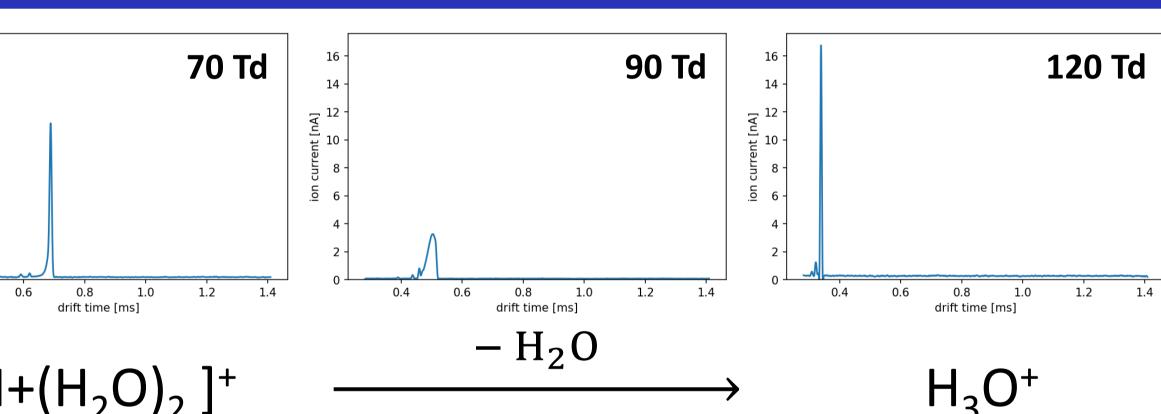


The reduced ion mobility can be calculated from the drift time. Cluster transitions can be clearly observed in the ion mobilities and intensities. By increasing water mixing ratio, the red. ion mobility is decreasing (increasing average cluster size, fig. 4).

# Conclusion

- Dynamic cluster systems at high red. field strength can be investigated with a "simple" analytical method.
- All cluster systems are dominated by water as neutral clustering agent, if there is not a high concentration of other neutral species.
- A qualitative rating of modifier effectiveness is possible.
- Cluster transitions can be detected. The transitions are generally influenced by humidity and temperature effects.
- result.
- Only a well running HiKE-IMS-MS-coupling can help to understand all signals (first steps are done).
- Findings can be transferred to other comparable systems like DMS or ion optics.
- The basic cluster reactions are confirmed by suitable simulations.





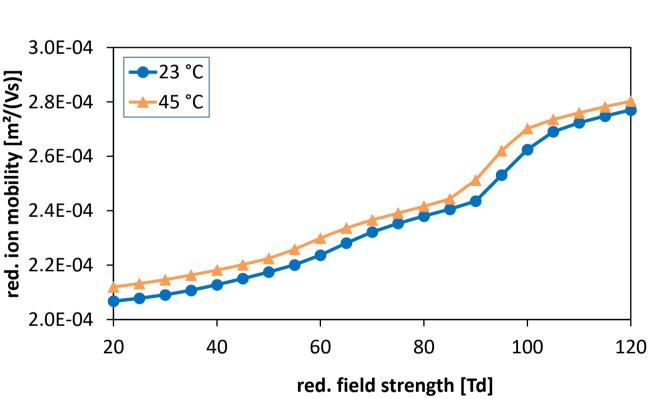


Figure 5: Effect of rising temperature in the drift tube on the red. ion mobility of water cluster system

Rising the temperature results in an increase of red. ion mobility (fig. 5). This can be explained by a lower average cluster size and a dynamic cluster system.

Figure 6: Cluster-transition processes at three representative reduced field strengths. The typical signals of the water-system (RIP) is marked with w and the analyte-specific signals with aX; (1.3 ppmV acetone in the reaction tube and 2.2 ppmV acetone in the drift gas)

Adding acetone results in three main analyte signals, which indicates multiple  $\frac{1}{2}$ <sup>2.0E-04</sup> cluster species. The humidity dependence of a1 and a2 indicates mixed wateranalyte clusters. The third signal seems to be the stable, non clustering acetone dimer (mobility known from previous results). The temperature effect confirms this notion: Note the decreasing cluster size by rising temperature. Modifying the drift gas does not show any effect with the used concentrations.

Even one analyte can generate a group of new species and a very complex chemical system in

# Outlook

- The modifier effect will continue to be investigated with HiKE-IMS, IMS and DMS: A selected set of particularly small analytes with typical functional groups will be measured with and without modifier.
- An In-depth investigation of the detected signals with a HiKE-IMS-MS-coupling is currently being carried out (Leibniz University Hannover).
- Kinetic simulations to explain the observed pre-RIP-signals are planned, as well as a further optimization of the used collision cross sections.

Experimental<sup>[5]</sup>

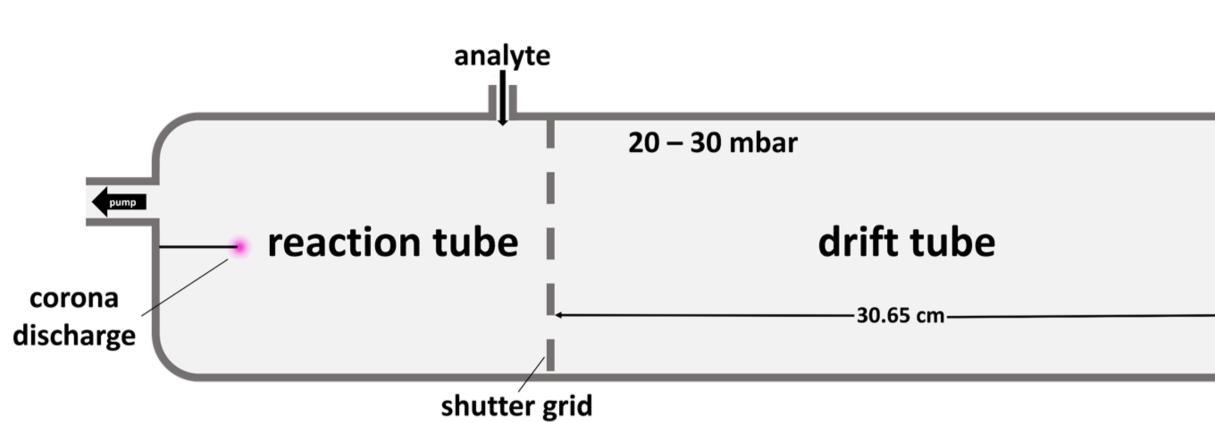
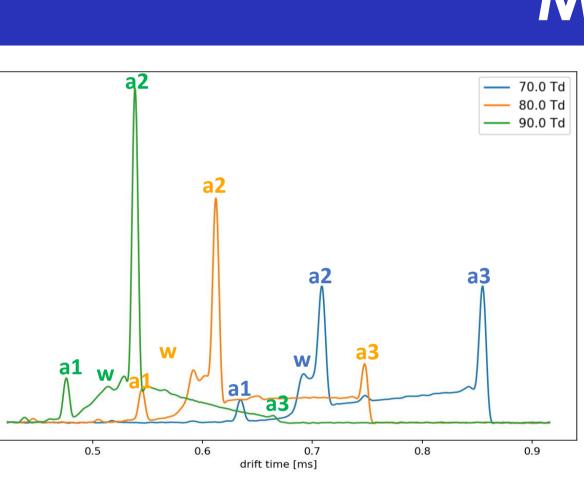
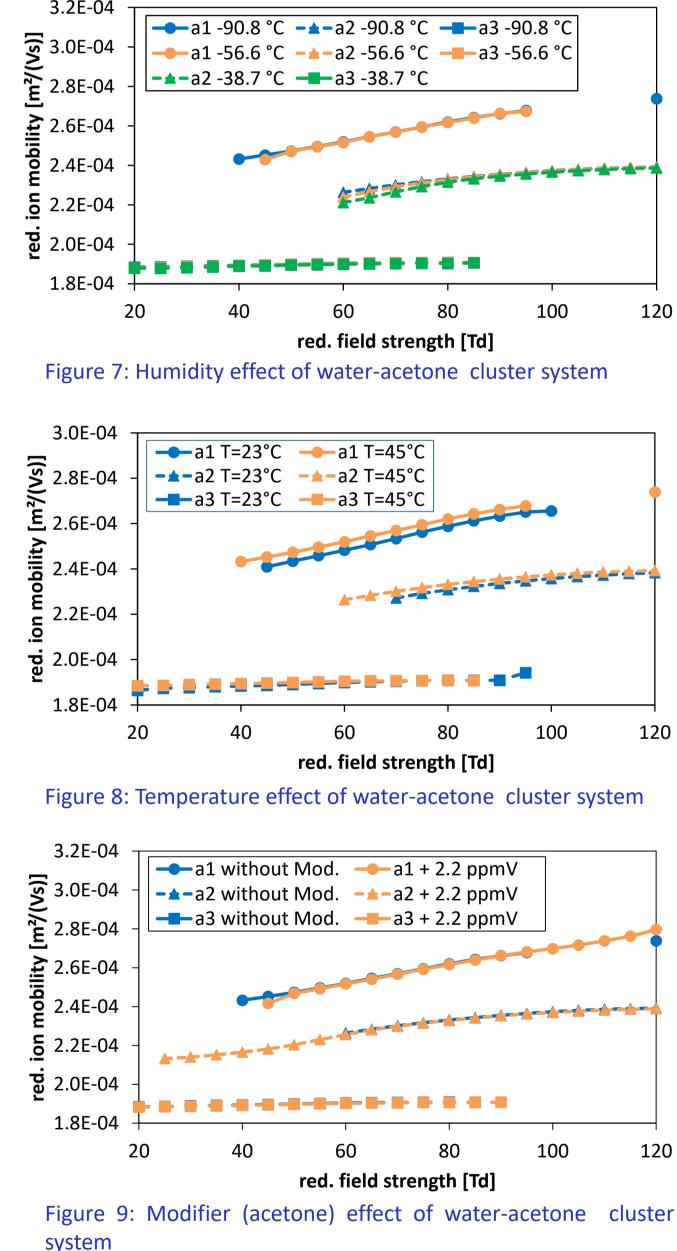


Figure 1: Schematic of the HiKE-IMS

### Simulations:

the chemical reaction system can be simulated by a particle based kinetic code (customized version of a Monte Carlo method)<sup>[6]</sup>





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  - Acknowledgement

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# Mixed water-acetone clusters



<sup>1</sup>: Physical & Theoretical Chemistry Wuppertal, Germany (Institute for Pure and Applied Mass Spectrometry)

<sup>2:</sup> Leibniz University Hannover

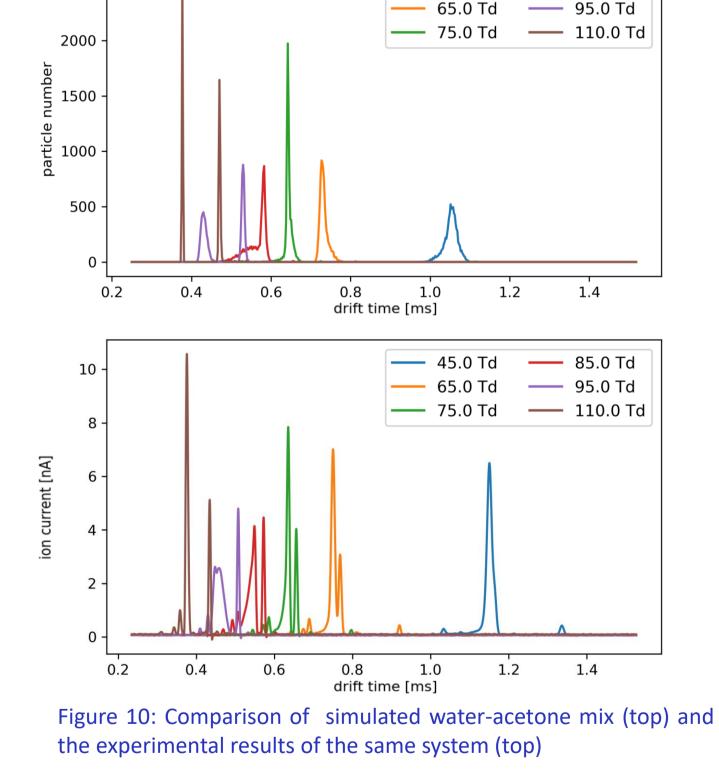
Institute of Electrical Engineering and Measurement Technology

### **HiKE-IMS**:

- basic setup close to a classic drift tube IMS
- pressure of 20-30 mbar in the reaction and the drift tube (30.65 cm)
- operated with pure nitrogen
- ionization by a corona discharge
- analytes are added in pure nitrogen to the reaction tube and the modifiable matrix drift gas to the drift tube
- heatable up to 45 °C

drift gas

The qualitative results can be reproduced by simulations (fig. 10): Just small shifts of all peaks, probably due to slightly inaccurate collision cross sections can be recognized. Note the cluster transition in both spectra at nearly the same red. field strength. 45.0 Td — 85.0 Td — 95.0 Td — 75.0 Td — 110.0 Td



## Literature

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